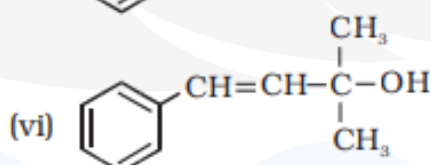
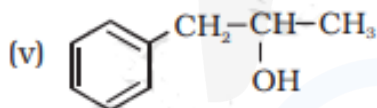
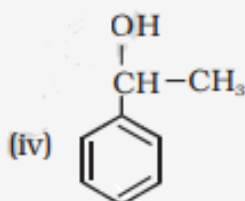
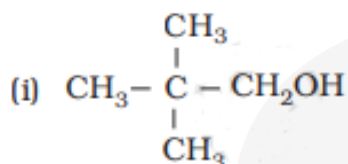


Chapter 7 - Alcohols, Phenols and Ethers XII

Intext Questions

Question 7.1:

Classify the following as primary, secondary and tertiary alcohols:



Answer

Primary alcohol \rightarrow (i), (ii), (iii)

Secondary alcohol \rightarrow (iv), (v)

Tertiary alcohol \rightarrow (vi)

Question 7.2:

Identify allylic alcohols in the above examples.

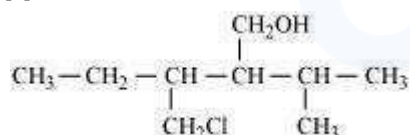
Answer

The alcohols given in (ii) and (vi) are allylic alcohols.

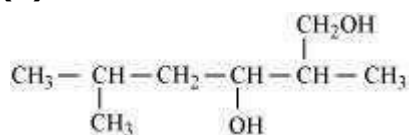
Question 7.3:

Name the following compounds according to IUPAC system.

(i)

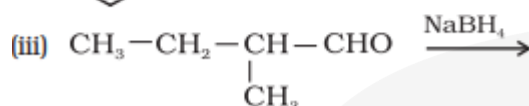
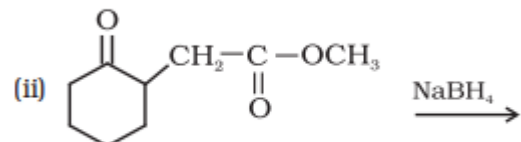
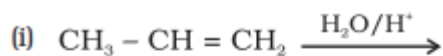


(ii)



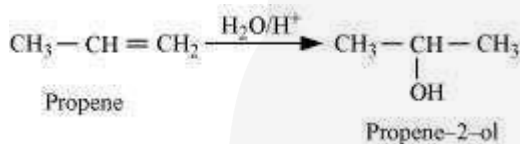
Question 7.5:

Write structures of the products of the following reactions:

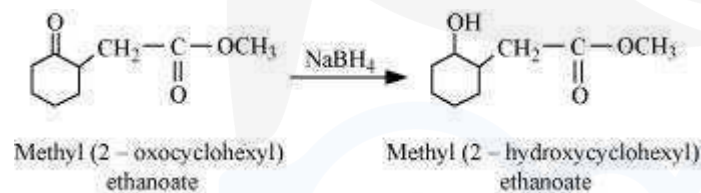


Answer

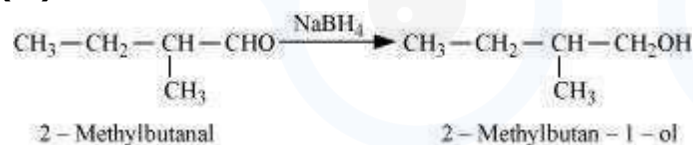
(i)



(ii)



(iii)

**Question 7.6:**

Give structures of the products you would expect when each of the following alcohol reacts with (a) HCl-ZnCl₂ (b) HBr and (c) SOCl₂.

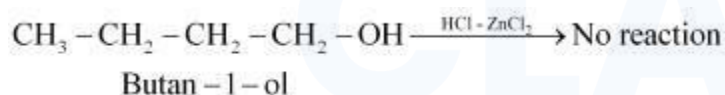
(i) Butan-1-ol

(ii) 2-Methylbutan-2-ol

Answer

(a)

(i)

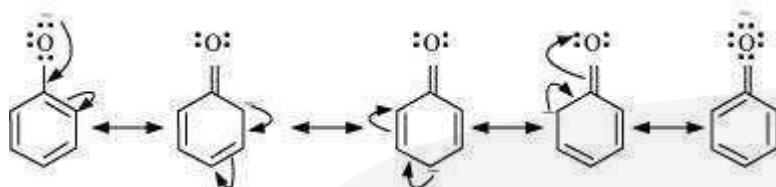


Primary alcohols do not react appreciably with Lucas' reagent (HCl-ZnCl₂) at room temperature.

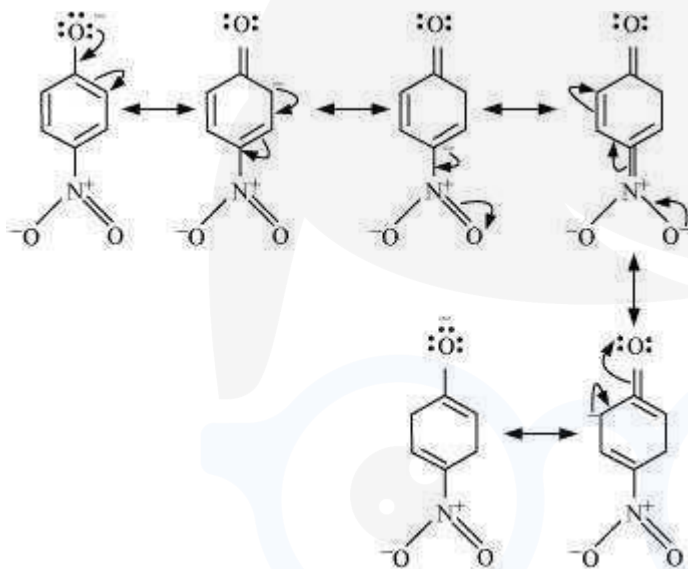
Question 7.8:

Ortho and *para* nitrophenols are more acidic than phenol. Draw the resonance structures of the corresponding phenoxide ions.

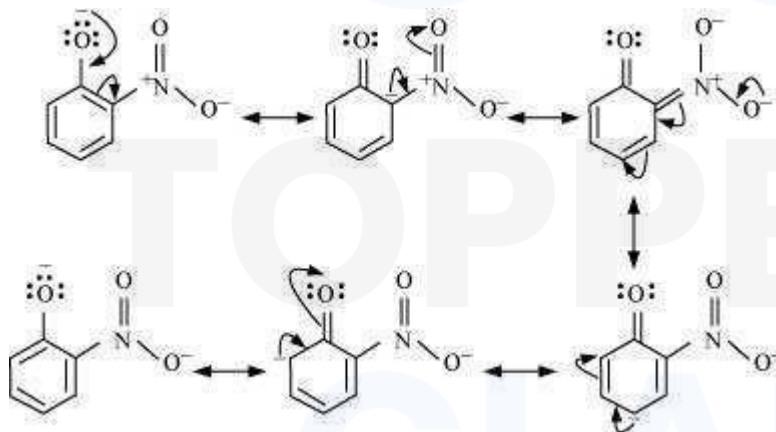
Answer



Resonance structure of the phenoxide ion



Resonance structures of *p*-nitrophenoxide ion



Resonance structures of *m*-nitrophenoxide ion

It can be observed that the presence of nitro groups increases the stability of phenoxide ion.

Question 7.9:

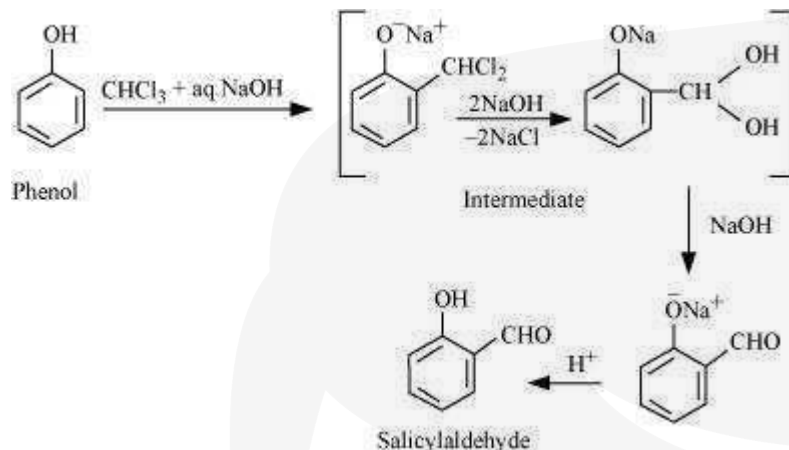
Write the equations involved in the following reactions:

(i) Reimer-Tiemann reaction

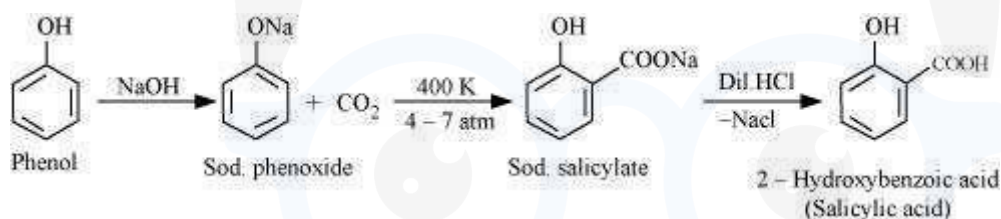
(ii) Kolbe's reaction

Answer

i. Reimer-Tiemann reaction



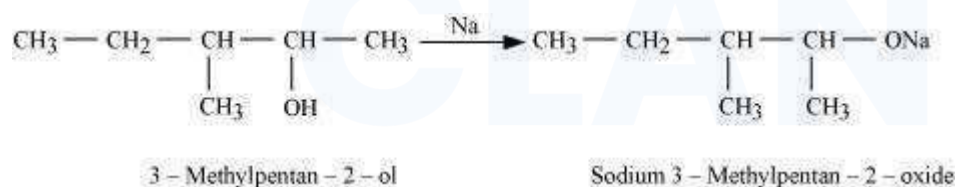
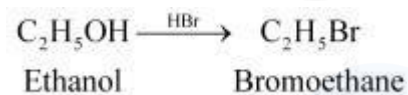
ii. Kolbe's reaction

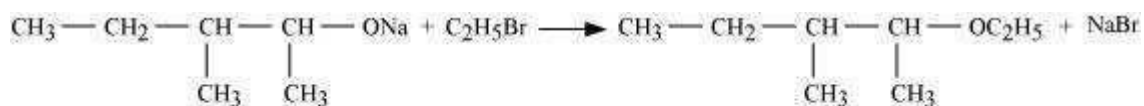
**Question 7.10:**

Write the reactions of Williamson synthesis of 2-ethoxy-3-methylpentane starting from ethanol and 3-methylpentan-2-ol.

Answer

In Williamson synthesis, an alkyl halide reacts with an alkoxide ion. Also, it is an $\text{S}_{\text{N}}2$ reaction. In the reaction, alkyl halides should be primary having the least steric hindrance. Hence, an alkyl halide is obtained from ethanol and alkoxide ion from 3-methylpentan-2-ol.





2 - Ethoxy - 3 - methylpentane

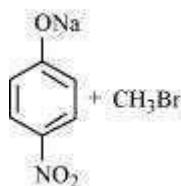
Question 7.11:

Which of the following is an appropriate set of reactants for the preparation of 1-methoxy-4-nitrobenzene and why?

(i)

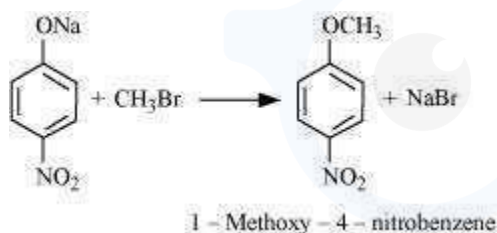


(ii)



Answer

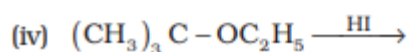
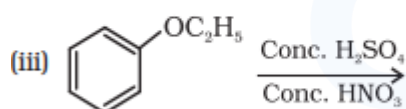
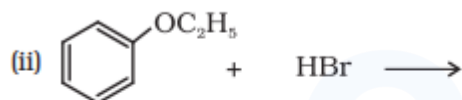
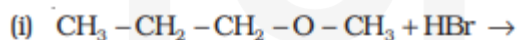
Set (ii) is an appropriate set of reactants for the preparation of 1-methoxy-4-nitrobenzene.



In set (i), sodium methoxide (CH_3ONa) is a strong nucleophile as well as a strong base. Hence, an elimination reaction predominates over a substitution reaction.

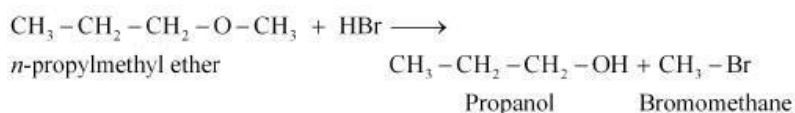
Question 7.12:

Predict the products of the following reactions:

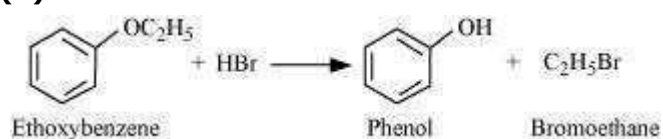


Answer

(i)



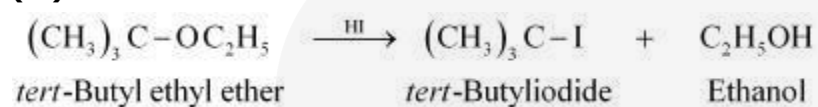
(ii)



(iii)



(iv)



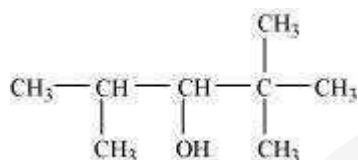
TOPPERS
CLAN

Exercises

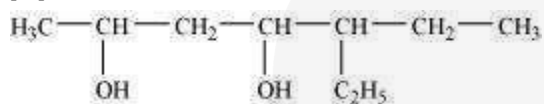
Question 7.1:

Write IUPAC names of the following compounds:

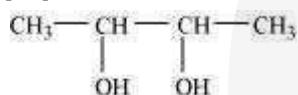
(i)



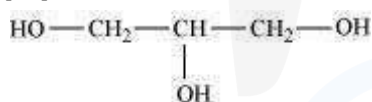
(ii)



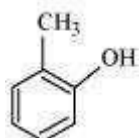
(iii)



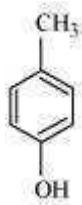
(iv)



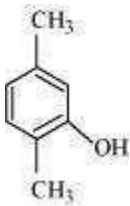
(v)



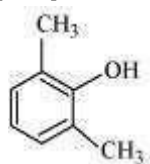
(vi)



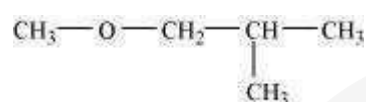
(vii)



(viii)



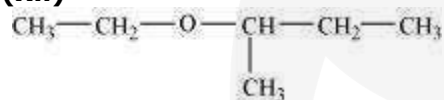
(ix)



(x) $\text{C}_6\text{H}_5-\text{O}-\text{C}_2\text{H}_5$

(xi) $\text{C}_6\text{H}_5-\text{O}-\text{C}_7\text{H}_{15}(\text{n-})$

(xii)



Answer

(i) 2, 2, 4-Trimethylpentan-3-ol

(ii) 5-Ethylheptane-2, 4-diol

(iii) Butane-2, 3-diol

(iv) Propane-1, 2, 3-triol

(v) 2-Methylphenol

(vi) 4-Methylphenol

(vii) 2, 5-Dimethylphenol

(viii) 2, 6-Dimethylphenol

(ix) 1-Methoxy-2-methylpropane

(x) Ethoxybenzene

(xi) 1-Phenoxyheptane

(xii) 2-Ethoxybutane

TOPPERS
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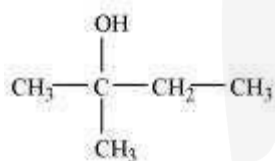
Question 7.2:

Write structures of the compounds whose IUPAC names are as follows:

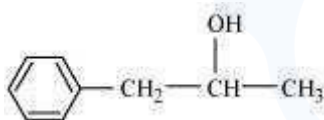
- (i) 2-Methylbutan-2-ol
- (ii) 1-Phenylpropan-2-ol
- (iii) 3,5-Dimethylhexane –1, 3, 5-triol
- (iv) 2,3 – Diethylphenol
- (v) 1 – Ethoxypropane
- (vi) 2-Ethoxy-3-methylpentane
- (vii) Cyclohexylmethanol
- (viii) 3-Cyclohexylpentan-3-ol
- (ix) Cyclopent-3-en-1-ol
- (x) 4-Chloro-3-ethylbutan-1-ol.

Answer

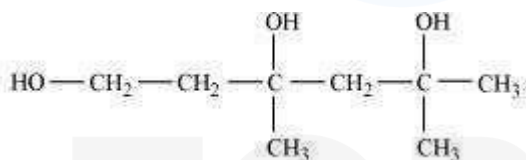
(i)



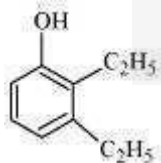
(ii)



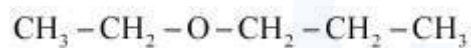
(iii)



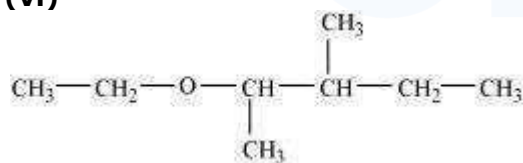
(iv)



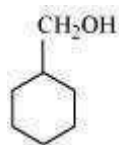
(v)



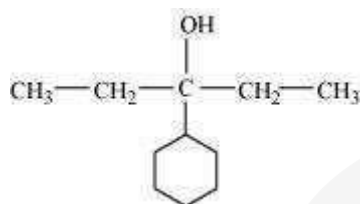
(vi)



(vii)



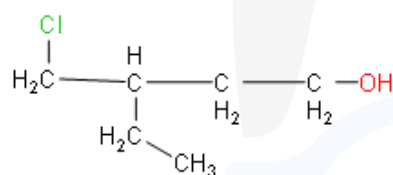
(viii)



(ix)



(x)



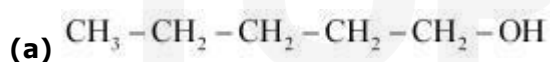
Question 7.3:

(i) Draw the structures of all isomeric alcohols of molecular formula $C_5H_{12}O$ and give their IUPAC names.

(ii) Classify the isomers of alcohols in question 11.3 (i) as primary, secondary and tertiary alcohols.

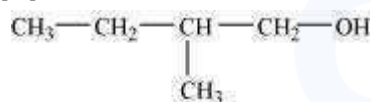
Answer

(i) The structures of all isomeric alcohols of molecular formula, $C_5H_{12}O$ are shown below:



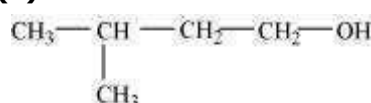
Pentan-1-ol (1°)

(b)



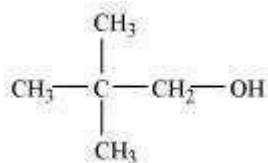
2-Methylbutan-1-ol (1°)

(c)



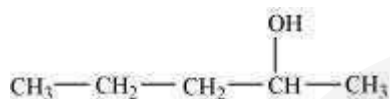
3-Methylbutan-1-ol (1°)

(d)



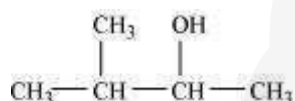
2, 2-Dimethylpropan-1-ol (1°)

(e)



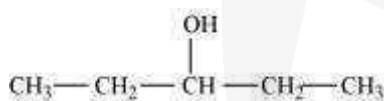
Pentan-2-ol (2°)

(f)



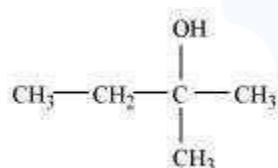
3-Methylbutan-2-ol (2°)

(g)



Pentan-3-ol (2°)

(h)



2-Methylbutan-2-ol (3°)

(ii) Primary alcohol: Pentan-1-ol; 2-Methylbutan-1-ol;

3-Methylbutan-1-ol; 2, 2-Dimethylpropan-1-ol

Secondary alcohol: Pentan-2-ol; 3-Methylbutan-2-ol;

Pentan-3-ol

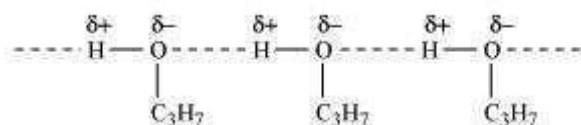
Tertiary alcohol: 2-methylbutan-2-ol

Question 7.4:

Explain why propanol has higher boiling point than that of the hydrocarbon, butane?

Answer

Propanol undergoes intermolecular H-bonding because of the presence of $-\text{OH}$ group. On the other hand, butane does not



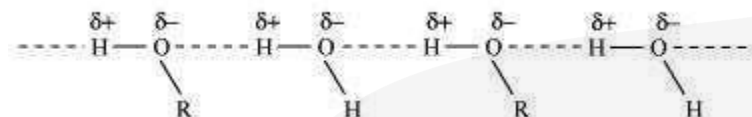
Therefore, extra energy is required to break hydrogen bonds. For this reason, propanol has a higher boiling point than hydrocarbon butane.

Question 7.5:

Alcohols are comparatively more soluble in water than hydrocarbons of comparable molecular masses. Explain this fact.

Answer

Alcohols form H-bonds with water due to the presence of $-OH$ group. However, hydrocarbons cannot form H-bonds with water.



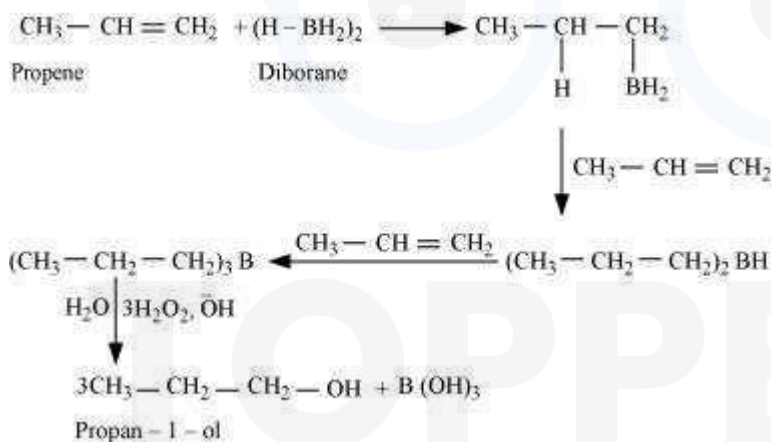
As a result, alcohols are comparatively more soluble in water than hydrocarbons of comparable molecular masses.

Question 7.6:

What is meant by hydroboration-oxidation reaction? Illustrate it with an example.

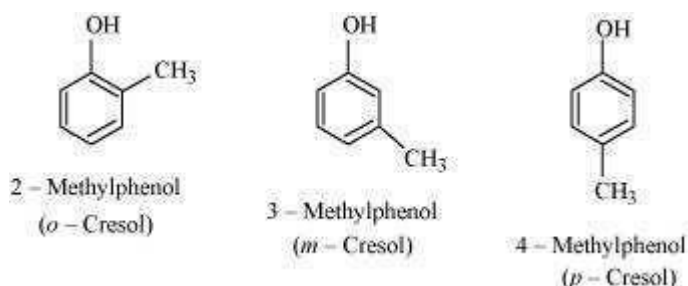
Answer

The addition of borane followed by oxidation is known as the hydroboration-oxidation reaction. For example, propan-1-ol is produced by the hydroboration-oxidation reaction of propene. In this reaction, propene reacts with diborane (BH_3)₂ to form trialkyl borane as an addition product. This addition product is oxidized to alcohol by hydrogen peroxide in the presence of aqueous sodium hydroxide.

**Question 7.7:**

Give the structures and IUPAC names of monohydric phenols of molecular formula, C_7H_8O .

Answer

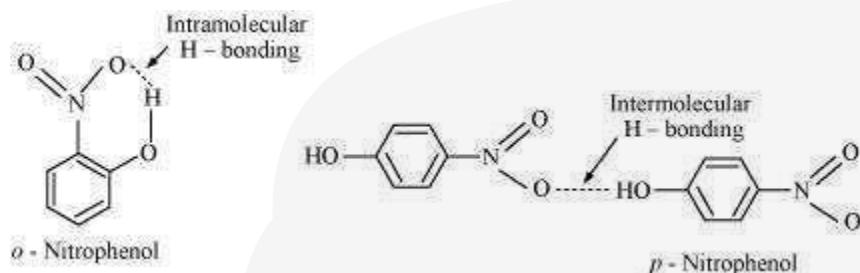


Question 7.8:

While separating a mixture of *ortho* and *para* nitrophenols by steam distillation, name the isomer which will be steam volatile. Give reason.

Answer

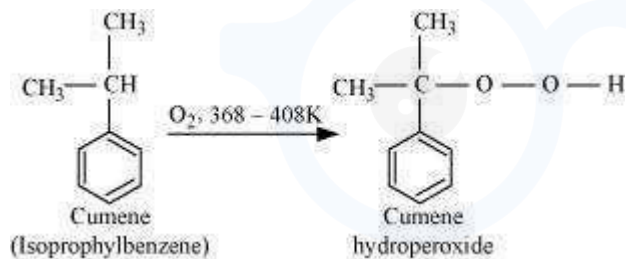
Intramolecular H-bonding is present in *o*-nitrophenol and *p*-nitrophenol. In *p*-nitrophenol, the molecules are strongly associated due to the presence of intermolecular bonding. Hence, *o*-nitrophenol is steam volatile.

**Question 7.9:**

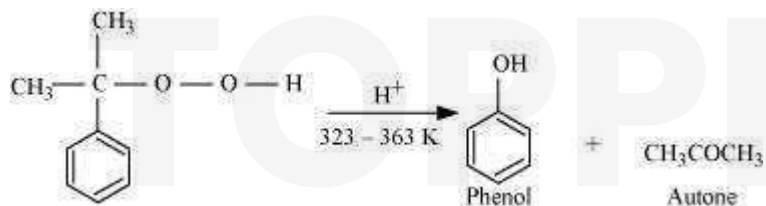
Give the equations of reactions for the preparation of phenol from cumene.

Answer

To prepare phenol, cumene is first oxidized in the presence of air of cumene hydroperoxide.



Then, cumene hydroxide is treated with dilute acid to prepare phenol and acetone as byproducts.

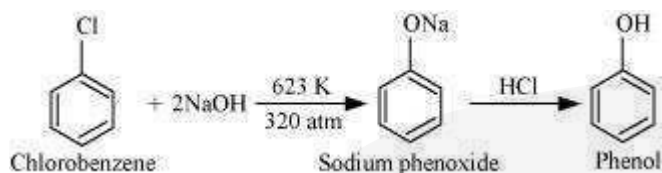


Question 7.10:

Write chemical reaction for the preparation of phenol from chlorobenzene.

Answer

Chlorobenzene is fused with NaOH (at 623 K and 320 atm pressure) to produce sodium phenoxide, which gives phenol on acidification.

**Question 7.11:**

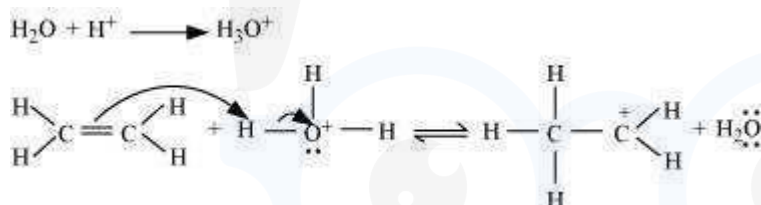
Write the mechanism of hydration of ethene to yield ethanol.

Answer

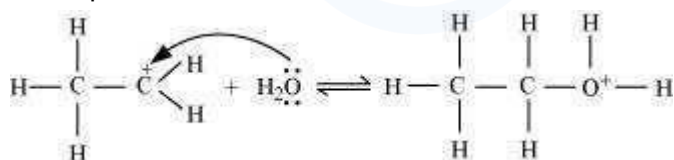
The mechanism of hydration of ethene to form ethanol involves three steps.

Step 1:

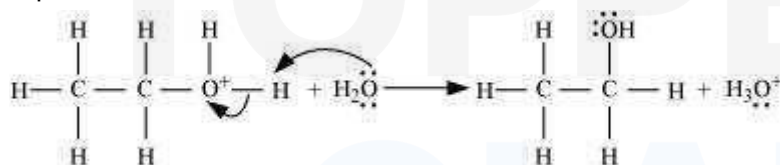
Protonation of ethene to form carbocation by electrophilic attack of H_3O^+ :

**Step 2:**

Nucleophilic attack of water on carbocation:

**Step 3:**

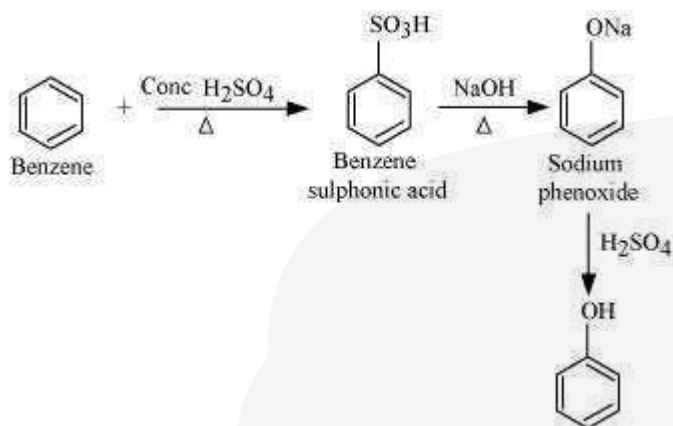
Deprotonation to form ethanol:



Question 7.12:

You are given benzene, conc. H_2SO_4 and NaOH . Write the equations for the preparation of phenol using these reagents.

Answer

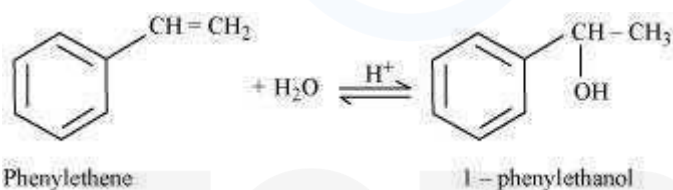
**Question 7.13:**

Show how will you synthesize:

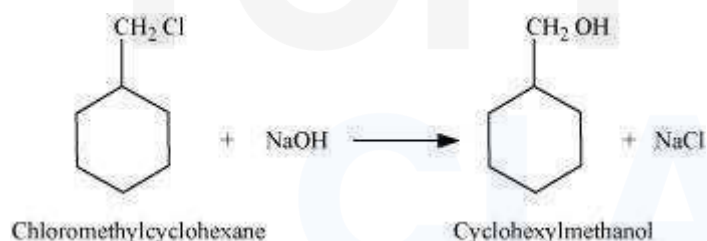
- (i) 1-phenylethanol from a suitable alkene.
- (ii) cyclohexylmethanol using an alkyl halide by an $\text{S}_{\text{N}}2$ reaction.
- (iii) pentan-1-ol using a suitable alkyl halide?

Answer

- (i) By acid-catalyzed hydration of ethylbenzene (styrene), 1-phenylethanol can be synthesized.



- (ii) When chloromethylcyclohexane is treated with sodium hydroxide, cyclohexylmethanol is obtained.



- (iii) When 1-chloropentane is treated with NaOH , pentan-1-ol is produced.



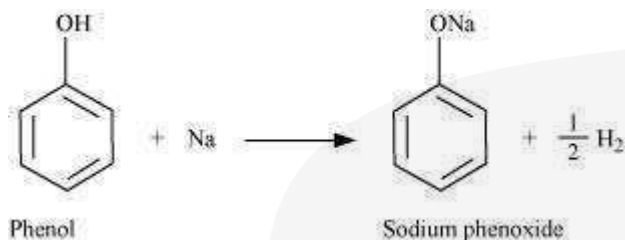
Question 7.14:

Give two reactions that show the acidic nature of phenol. Compare acidity of phenol with that of ethanol.

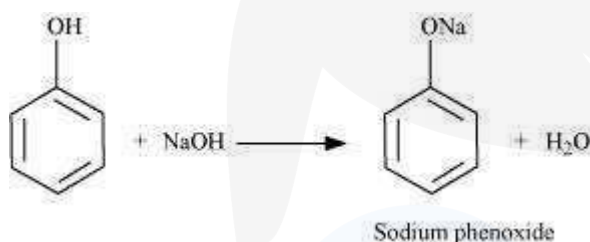
Answer

The acidic nature of phenol can be represented by the following two reactions:

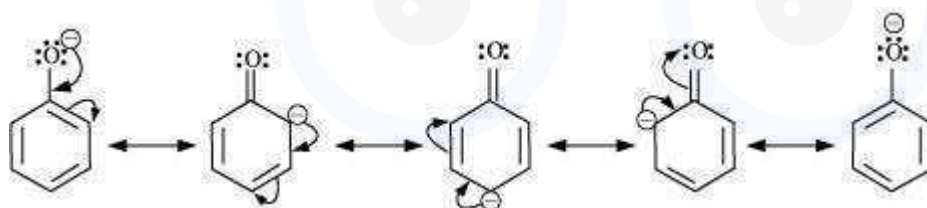
(i) Phenol reacts with sodium to give sodium phenoxide, liberating H₂.



(ii) Phenol reacts with sodium hydroxide to give sodium phenoxide and water as by-products.

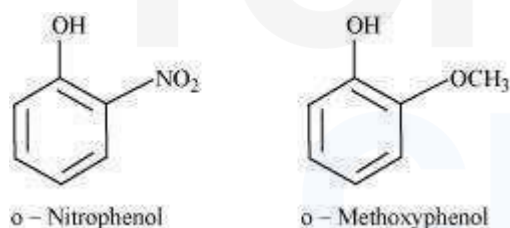


The acidity of phenol is more than that of ethanol. This is because after losing a proton, the phenoxide ion undergoes resonance and gets stabilized whereas ethoxide ion does not.

**Question 7.15:**

Explain why is *ortho* nitrophenol more acidic than *ortho* methoxyphenol?

Answer



The nitro-group is an electron-withdrawing group. The presence of this group in the ortho position decreases the electron density in the O–H bond. As a result, it is easier to lose a proton. Also, the *o*-nitrophenoxide ion formed after the loss of protons is stabilized by resonance. Hence, *ortho* nitrophenol is a stronger acid.

On the other hand, methoxy group is an electron-releasing group. Thus, it increases the electron density in the O–H bond and hence, the proton cannot be given out easily.

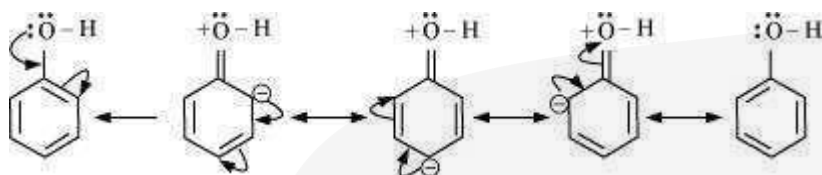
For this reason, *ortho*-nitrophenol is more acidic than *ortho*-methoxyphenol.

Question 7.16:

Explain how does the –OH group attached to a carbon of benzene ring activate it towards electrophilic substitution?

Answer

The –OH group is an electron-donating group. Thus, it increases the electron density in the benzene ring as shown in the given resonance structure of phenol.



As a result, the benzene ring is activated towards electrophilic substitution.

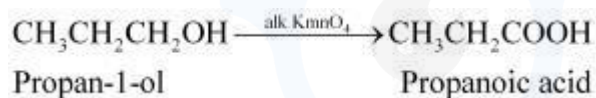
Question 7.17:

Give equations of the following reactions:

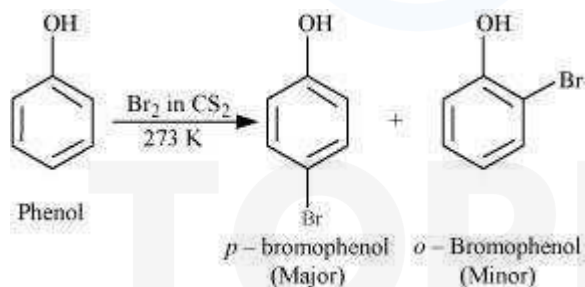
- (i) Oxidation of propan-1-ol with alkaline KMnO_4 solution.
- (ii) Bromine in CS_2 with phenol.
- (iii) Dilute HNO_3 with phenol.
- (iv) Treating phenol with chloroform in presence of aqueous NaOH .

Answer

(i)



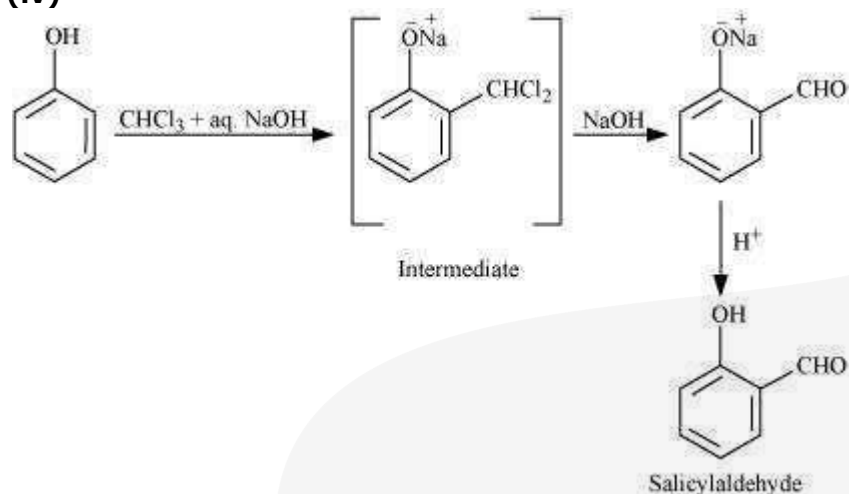
(ii)



(iii)



(iv)



Question 7.18:

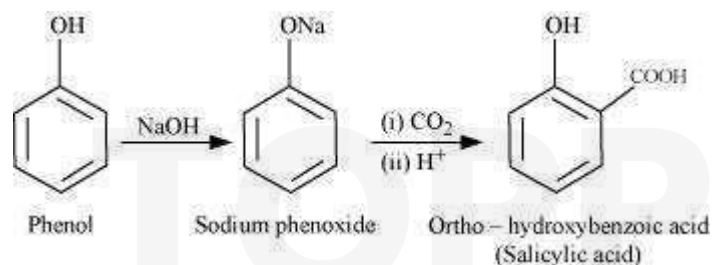
Explain the following with an example.

- (i) Kolbe's reaction.
- (ii) Reimer-Tiemann reaction.
- (iii) Williamson ether synthesis.
- (iv) Unsymmetrical ether.

Answer

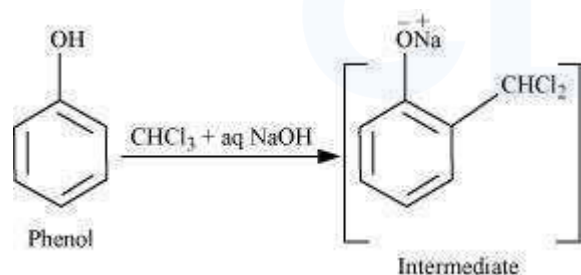
(i) Kolbe's reaction:

When phenol is treated with sodium hydroxide, sodium phenoxide is produced. This sodium phenoxide when treated with carbon dioxide, followed by acidification, undergoes electrophilic substitution to give ortho-hydroxybenzoic acid as the main product. This reaction is known as Kolbe's reaction.



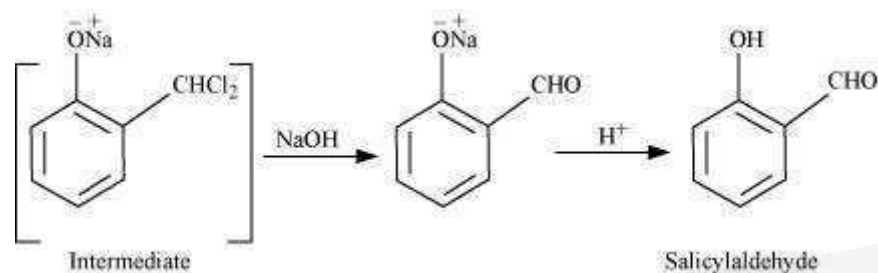
(ii) Reimer-Tiemann reaction:

When phenol is treated with chloroform (CHCl_3) in the presence of sodium hydroxide, a $-\text{CHO}$ group is introduced at the ortho position of the benzene ring.



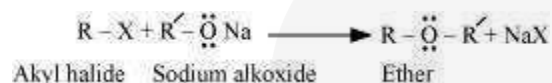
This reaction is known as the Reimer-Tiemann reaction.

The intermediate is hydrolyzed in the presence of alkalis to produce salicylaldehyde.



(iii) Williamson ether synthesis:

Williamson ether synthesis is a laboratory method to prepare symmetrical and unsymmetrical ethers by allowing alkyl halides to react with sodium alkoxides.



This reaction involves $\text{S}_{\text{N}}2$ attack of the alkoxide ion on the alkyl halide. Better results are obtained in case of primary alkyl halides.



If the alkyl halide is secondary or tertiary, then elimination competes over substitution.

(iv) Unsymmetrical ether:

An unsymmetrical ether is an ether where two groups on the two sides of an oxygen atom differ (i.e., have an unequal number of carbon atoms). For example: ethyl methyl ether ($\text{CH}_3\text{-O-CH}_2\text{CH}_3$).

Question 7.19:

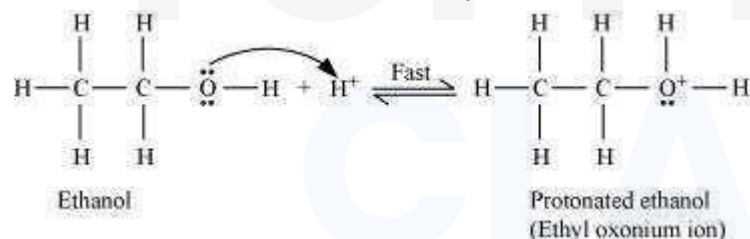
Write the mechanism of acid dehydration of ethanol to yield ethene.

Answer

The mechanism of acid dehydration of ethanol to yield ethene involves the following three steps:

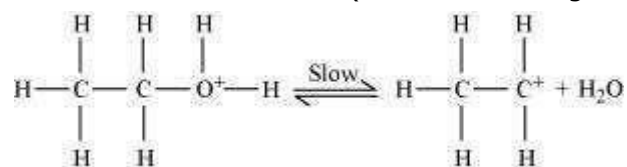
Step 1:

Protonation of ethanol to form ethyl oxonium ion:



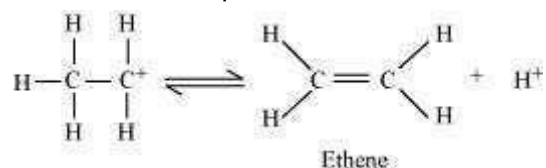
Step 2:

Formation of carbocation (rate determining step):



Step 3:

Elimination of a proton to form ethene:



The acid consumed in step 1 is released in Step 3. After the formation of ethene, it is removed to shift the equilibrium in a forward direction.

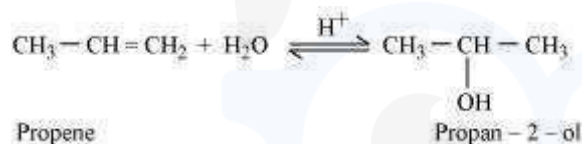
Question 7.20:

How are the following conversions carried out?

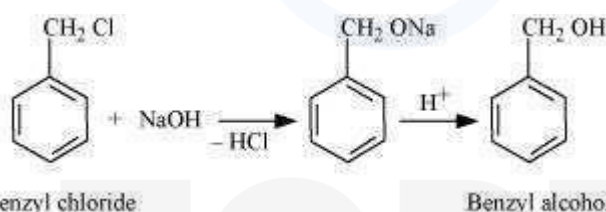
- (i) Propene \rightarrow Propan-2-ol
- (ii) Benzyl chloride \rightarrow Benzyl alcohol
- (iii) Ethyl magnesium chloride \rightarrow Propan-1-ol.
- (iv) Methyl magnesium bromide \rightarrow 2-Methylpropan-2-ol.

Answer

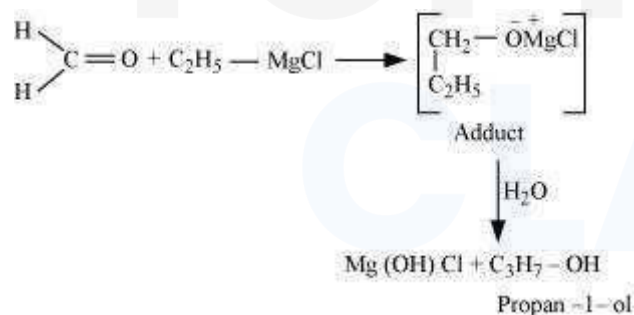
(i) If propene is allowed to react with water in the presence of an acid as a catalyst, then propan-2-ol is obtained.



(ii) If benzyl chloride is treated with NaOH (followed by acidification) then benzyl alcohol is produced.



(iii) When ethyl magnesium chloride is treated with methanal, an adduct is the product which gives propan-1-ol on hydrolysis.

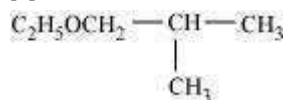


(iv) When methyl magnesium bromide is treated with propane, an adduct is the product which gives 2-methylpropan-2-ol on hydrolysis.

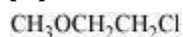
Question 7.23:

Give IUPAC names of the following ethers:

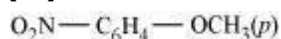
(i)



(ii)



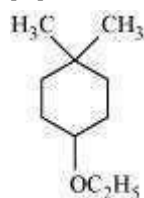
(iii)



(iv)



(v)



(vi)



Answer

(i) 1-Ethoxy-2-methylpropane

(ii) 2-Chloro-1-methoxyethane

(iii) 4-Nitroanisole

(iv) 1-Methoxypropane

(v) 1-Ethoxy-4,4-dimethylcyclohexane

(vi) Ethoxybenzene

Question 7.24:

Write the names of reagents and equations for the preparation of the following ethers by Williamson's synthesis:

(i) 1-Propoxypropane

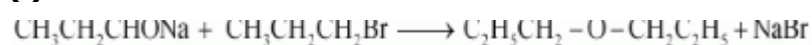
(ii) Ethoxybenzene

(iii) 2-Methoxy-2-methylpropane

(iv) 1-Methoxyethane

Answer

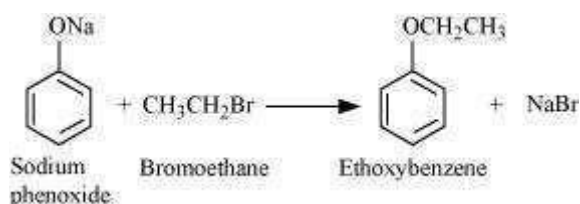
(i)



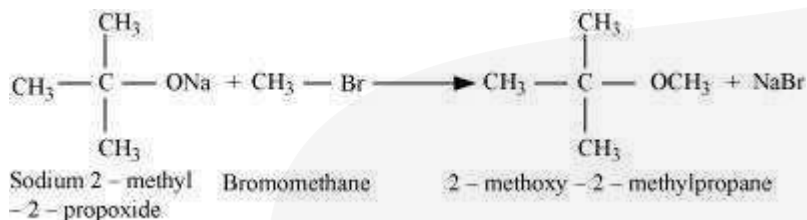
Sodium propoxide 1-Bromopropane

1-Propoxypropane

(ii)



(iii)



(iv)

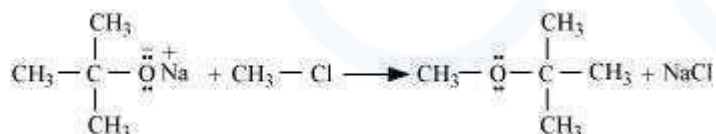


Question 7.25:

Illustrate with examples the limitations of Williamson synthesis for the preparation of certain types of ethers.

Answer

The reaction of Williamson synthesis involves $\text{S}_{\text{N}}2$ attack of an alkoxide ion on a primary alkyl halide.



But if secondary or tertiary alkyl halides are taken in place of primary alkyl halides, then elimination would compete over substitution. As a result, alkenes would be produced. This is because alkoxides are nucleophiles as well as strong bases. Hence, they react with alkyl halides, which results in an elimination reaction.



Tertiary alkyl halide

Alkene

Question 7.26:

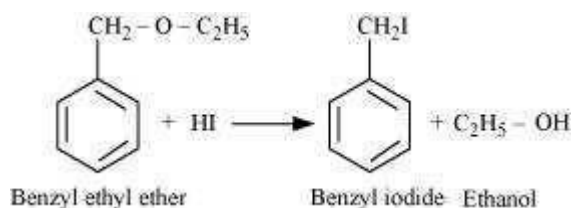
How is 1-propoxypropane synthesised from propan-1-ol? Write mechanism of this reaction.

Answer

1-propoxypropane can be synthesized from propan-1-ol by dehydration.

Propan-1-ol undergoes dehydration in the presence of protic acids (such as H_2SO_4 , H_3PO_4) to give 1-propoxypropane.

(iii)



Question 7.29:

Explain the fact that in aryl alkyl ethers

(i) The alkoxy group activates the benzene ring towards electrophilic substitution and

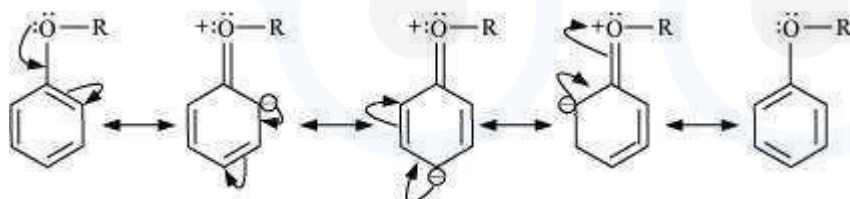
(ii) It directs the incoming substituents to ortho and para positions in benzene ring.

Answer

(i)



In aryl alkyl ethers, due to the +R effect of the alkoxy group, the electron density in the benzene ring increases as shown in the following resonance structure.



Thus, benzene is activated towards electrophilic substitution by the alkoxy group.

(ii) It can also be observed from the resonance structures that the electron density increases more at the ortho and para positions than at the meta position. As a result, the incoming substituents are directed to the ortho and para positions in the benzene ring.

Question 7.30:

Write the mechanism of the reaction of HI with methoxymethane.

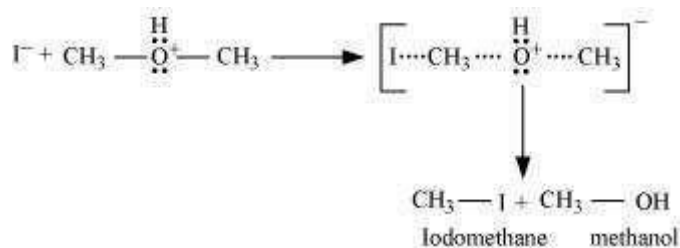
Answer

The mechanism of the reaction of HI with methoxymethane involves the following steps:

Step1: Protonation of methoxymethane:

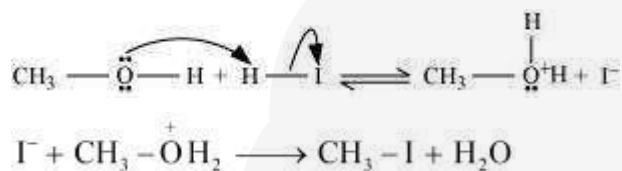


Step2: Nucleophilic attack of I⁻:



Step3:

When HI is in excess and the reaction is carried out at a high temperature, the methanol formed in the second step reacts with another HI molecule and gets converted to methyl iodide

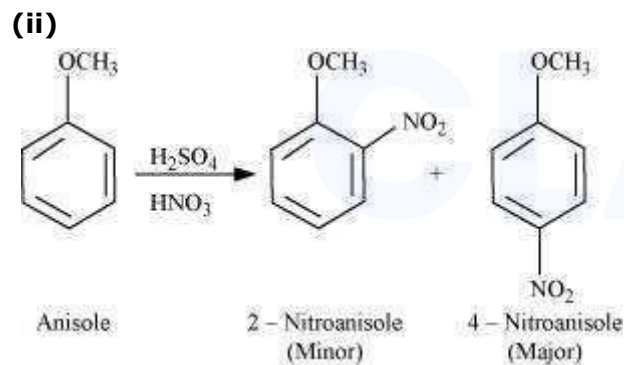
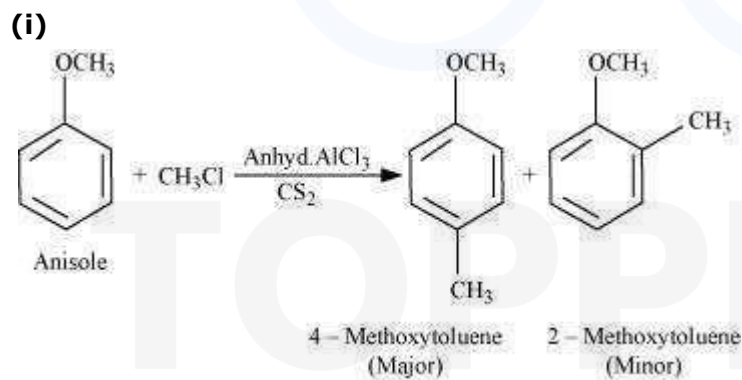


Question 7.31:

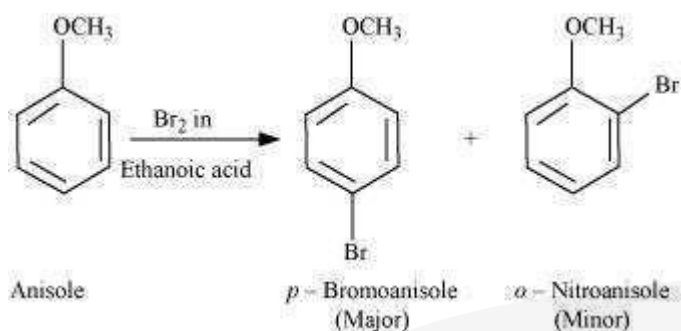
Write equations of the following reactions:

- (i) Friedel-Crafts reaction—alkylation of anisole.
- (ii) Nitration of anisole.
- (iii) Bromination of anisole in ethanoic acid medium.
- (iv) Friedel-Craft's acetylation of anisole.

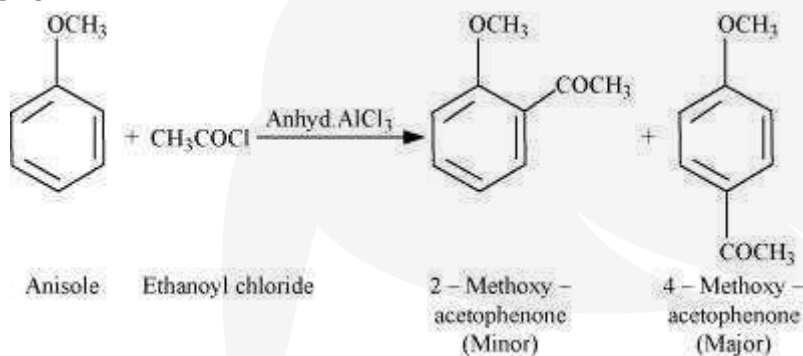
Answer



(iii)



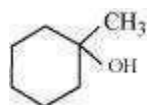
(iv)



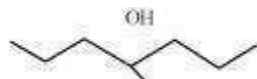
Question 7.32:

Show how would you synthesise the following alcohols from appropriate alkenes?

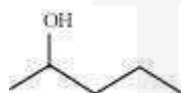
(i)



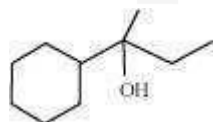
(ii)



(iii)



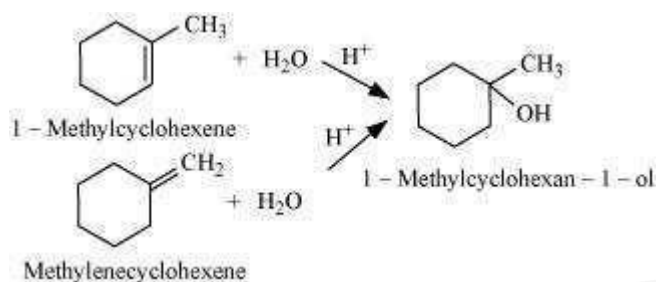
(iv)



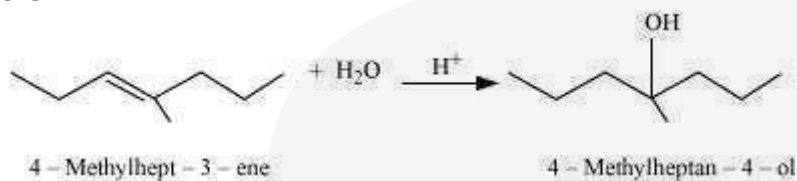
Answer

The given alcohols can be synthesized by applying Markovnikov's rule of acid-catalyzed hydration of appropriate alkenes.

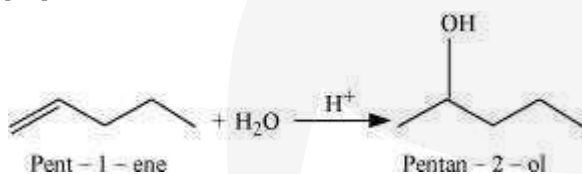
(i)



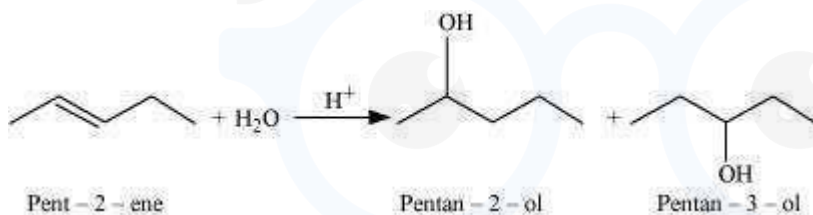
(ii)



(iii)

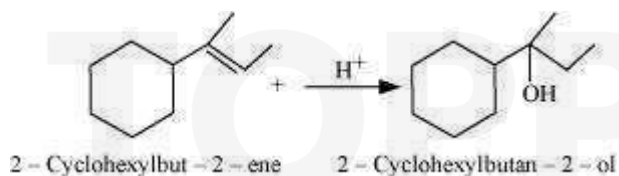


Acid-catalyzed hydration of pent-2-ene also produces pentan-2-ol but along with pentan-3-ol.



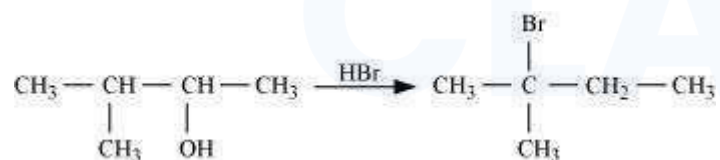
Thus, the first reaction is preferred over the second one to get pentan-2-ol.

(iv)



Question 7.33:

When 3-methylbutan-2-ol is treated with HBr, the following reaction takes place:



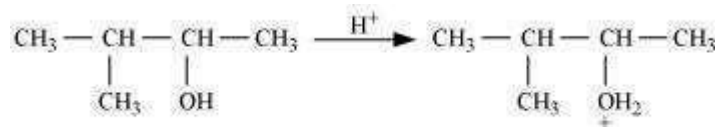
Give a mechanism for this reaction.

(Hint : The secondary carbocation formed in step II rearranges to a more stable tertiary carbocation by a hydride ion shift from 3rd carbon atom.)

Answer

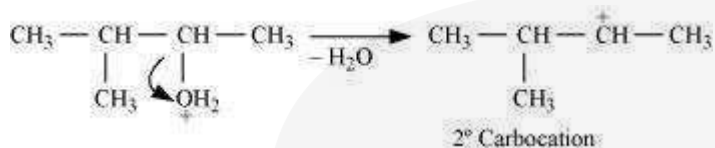
The mechanism of the given reaction involves the following steps:

Step 1: Protonation

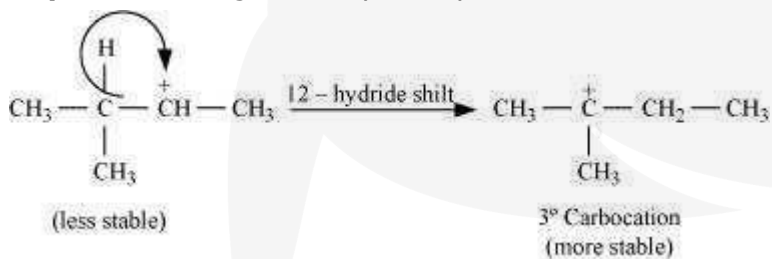


3 - Methylbutan - 2 - ol

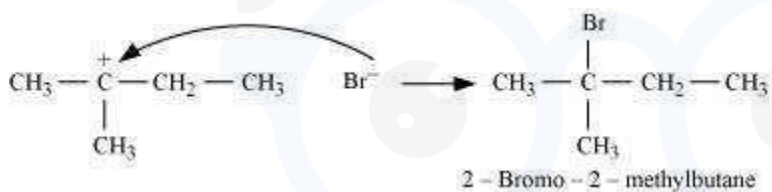
Step 2: Formation of 2° carbocation by the elimination of a water molecule



Step 3: Re-arrangement by the hydride-ion shift



Step 4: Nucleophilic attack



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